

Research Article

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Gallium (III)-Metalloporphyrin Grafted Magnetite Nanoparticles for Fluoride Removal from Aqueous Solutions

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Abstract

Fluoride is a critical micronutrient for human health and is usually ingested through food and drinking water. This technique has done in monitoring the concentration of fluoride in water complications at mining operation that it effects on the environment. Fe_3O_4 nanoparticles were synthesized by coprecipitation of Fe^{2+} and Fe^{3+} in an alkaline solution, their surface was then modified with 3-aminopropyltriethoxysilane and a Ga(III) porphyrin complex, Ga(TCPP)Cl₂[TCPP: tetrakis (4-carboxyphenyl) porphyrin]. The morphology and physical properties of these modified magnetic nanoparticles were examined by X-ray diffraction (XRD), Transmission electron microscopy (TEM), Fourier transform infrared spectroscopy (FT-IR), and thermogravimetric analysis (TGA). The batch experiments showed that the nanoparticles could be used to remove fluoride ions from aqueous solution, achieving an extraction efficiency in excess of 96% for an initial fluoride concentration of 10 mg/L. This was obtained at pH 5.5 using 50 mg of modified nanoparticles and a contact time of 30 min.

Keywords: Fluoride ion; Metalloporphyrin; Magnetite; Nanoparticles; Wastewater treatment; Surface modification

Introduction

Fluoride in the environment happens through its natural existence on the earth's crust and industrial activities, especially electroplating, glass, steel, aluminum, and enricher industries. The discharge of such wastewater leads to the fluoride contamination of surface and ground water. The contamination of ground water with fluoride can occur naturally by the dissolution of fluoride rich minerals whereas artificial pollution occurs due to the waste water produced from mine, metal refining, plating industries, aluminum smelters and high-tech semiconductor factories [1-5]. Fluoride pollution of the industrial wastewater is one of the significant problems worldwide imposing a severe threat to human health [6]. This method has done in controlling the concentration of fluoride in water problems at mining operation that it effects on the environment. Mine water emanates from working and abandoned surface and underground mines. Usually, it is ground water percolating through the host rocks or sediments, collected by the shafts, galleries, or open pits. Although mine water is not necessarily contaminated, in many cases it contains elevated metal or acidity values which cause a severe threat to the receiving water courses and river catchments. Noncontaminated mine water is used for several purposes, e.g., as cooling water for power plants, drinking water, mineral water, or for agricultural purposes [7]. As an environmental element, fluoride has received much attention due to its both useful and harmful effect on health. For example, while low concentration of fluoride in drinking water can cause dental cavities, excess concentration in drinking water cause dental and skeletal fluorosis, indicating the importance of standardization of this element at a recommended range in drinking water was 1.5 mg/L by World Health Organization (WHO). In order to reduce the concentration of fluoride in water, various defluoridation technologies have been developed,

including: viz., adsorption, ion exchange, precipitation and the membrane process such as reverse osmosis [8], Nano filtration [9], electro dialysis [10] and Donnan dialysis [11]. Adsorption is an efficient and common process in de-fluoridation of water, with satisfactory results and vast applications. Different sorbents have been investigated for removing fluoride, such as: active alumina [12], clay [13], activated carbon [14] and some other sorbents [15]. Recently, many studies have been concentrated on developing novel ionophores with useful specificity for anion [16]. For example, it is found that Ga-octaethyl porphyrin [OEP] [17,18] can be applied as a useful ionophore in producing membrane electrodes with significantly enhanced properties, such as favorable selectivity. In Ga(III)-metallopheyrin, the main reason for enhanced selectivity of fluoride can be attributed to the strong interaction between fluoride anion and the Ga(III) cation available in the structure of metalloporphyrin. As building blocks, nanoparticles play an important role in producing nanomaterials, due to their small size, which gives them a high surface area-to-volume ratio, resulting in new magnetic, electrical and optical properties. Improving chemical reactivity is another significant effect of the large surface area of nanoparticles, which provides the possibility for chemical interactions with other chemical species that are available in the environment. There are other abilities for using of nanoparticles such as: adsorption and recovery of some ligands, dyes and ionic species, which are available in the literature [19-24]. The main specification of these nanosorbents, and one which can be considered as their main advantage, is their ability to recover nanoparticles using their external magnetic field. In the mentioned material, dipole/dipole attraction and agglomeration can be created as a basis and a passive layer; this is attributable to important capabilities like high chemical activity, their oxidation ability in air and the loss of magnetism. However, some operations, such as modification of post- synthetic surface are essential [25-31]. For coated metal oxides at the nanoscale, the most common modification method is aminopropyltriethoxysilane (APTES). However, nanomaterials lack

the target-selective ability, which makes them a weak candidate for samples with complicated matrices. All of this indicates the necessity for functionalization as a tool for inducing selective targeting in nanomaterials [32-40]. For example, Ga-porphyrin is a useful ionophore which can be used in membrane electrodes [41,42]. In other words, it is possible to fabricate a new type of magnetic nanosorbent with high affinity toward fluoride and favorable magnetic reparability by combining the Ga(III) porphyrin and the magnetic nanoparticle. This paper investigates a new type of nanomagnet which is an APTES coated magnetite, functionalized with the Ga(III) tetra (4-carboxy phenyl) porphyrin dichloride are represented in Figure 1.

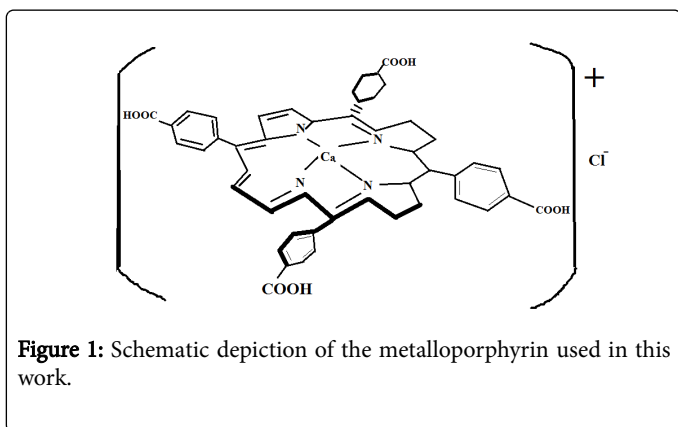


Figure 1: Schematic depiction of the metalloporphyrin used in this work.

Experimental

Materials

The most important materials were ferrous chloride tetrahydrate ($\text{FeCl}_2 \cdot 4\text{H}_2\text{O}$), dichloromethane (DCM), Ferric chloride hexahydrate ($\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$), APTES, methanol, ethanol, dicyclohexylcarbodiimide (DCHC), 4-paracarboxybenzaldehyde, propynoyic acid, pyrrole, N, N-dimethylformamide (DMF), gallium acetate, dimethyl sulfoxide (DMSO), sodium fluoride, sodium salts of indicated anions and de-ionized water in all experiments were graded analytically. Free basic tetrakis (4-carboxyphenyl) porphyrin (TCPP) (Aldrich) was dried under vacuum at 60 °C before using $\text{Ga}(\text{TCPP})\text{Cl}_2$ was prepared using methods described in the literature [43,44].

Equipment

Morphology of the samples was investigated using TEM with Zeiss 900 and 80 KV voltage. Powder X-ray diffraction (XRD) patterns of the solid produces in the 2θ range of 5-80° using a PW-1840 diffractometer were found from Philips Co. [45] with $\text{Cu } \alpha$ radiation ($\lambda = 1.54178 \text{ \AA}$). Thermogravimetric analysis was performed using a Mettler Toledo TGA/SDTA 851. Samples were heated under a static air atmosphere from ambient to 1000°C at a rate of 10°C/min. Ion chromatography (IC) determinations were carried out using a S1122 Sykam ion chromatograph. pH measurements were performed with a Denver Instrument Company 9710.1 pH meter kit UB-5. FT-IR spectra were recorded with a IFS 88 FT-IR spectrophotometer from Bruker Co., using a standard KBr pellet technique.

Sorbent preparation

Synthesis of Fe_3O_4 nanoparticles: Co-precipitation is known to be one of the best methods for synthesis of Fe_3O_4 nanoparticles, based on

a study performed by Maity [45,46]. Using this method, 2 g $\text{FeCl}_2 \cdot 4\text{H}_2\text{O}$ and 5.2 g $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$ were used, with 25 mL deionized water added. The mixed solution was stirred under N_2 at 80°C for 1 h. 100 mL of aqueous 1.5 M NaOH was then injected into the mixture and stirred under N_2 for 1 h. Next, precipitated particles were washed with deionized water three times in order to remove mineral salts; the solution was subsequently washed with ethanol three times to remove organic materials. In the next step precipitated particles were separated with magnets, and finally Fe_3O_4 nanoparticles were dried in an oven at 50°C for 24 h.

Modification of Fe_3O_4 nanoparticles with APTES group: Fe_3O_4 / APTES nanoparticles were synthesized according to a previously described method [47,48]. This reaction was between APTES and a hydroxyl group on the surface of the magnetite available in the alkaline solution; the bond Fe-O-H changed with Fe-O-Si [49]. 2.3 g magnetic nanoparticles, were dispersed in 100 mL ethanol using a sonication method, for about 1 h; under continuous mechanical stirring, 20.34 mL APTES was added. This reaction was achieved for 20 h under N_2 atmosphere. The Fe_3O_4 / APTES nanoparticles were prepared and separated with a magnet, and were washed with ethanol and deionized water three times. Finally, Fe_3O_4 / APTES nanoparticles were dried under vacuum at 70°C.

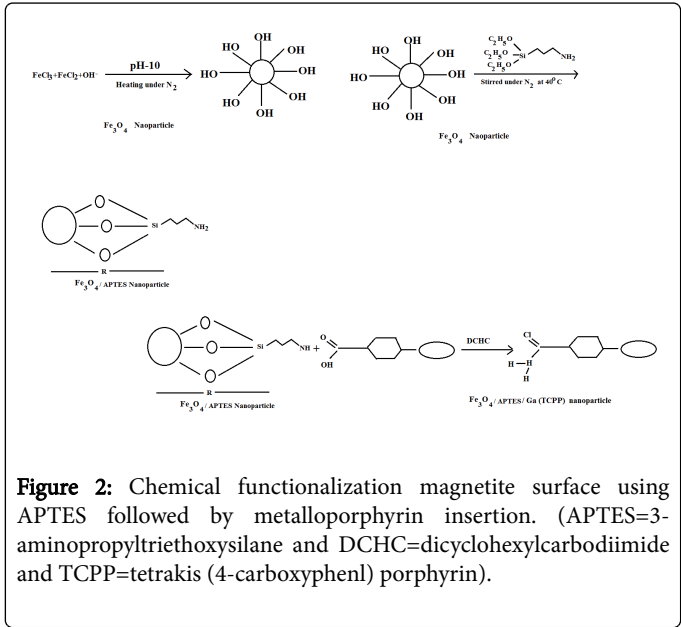
Grafting of Ga(TCPP) group at the surface of Fe_3O_4 / APTES: The metalloporphyrin structure is designed to accept different ions, based on central metal and complex formation; the metalloporphyrin is used instead of pulling electron groups, which gathers them on non-minerals support, making them a good choice for separation and specific to ions [50]. 10 g paracarboxide benzaldehyde and 5 mL pyrrole, mixed in the bottom flask, were stirred under N_2 , followed by the addition of 250 mL propionic acid. The mixture was refluxed at 140°C for 30 min. The product was filtered with filter papers that does not have disturbance in this reaction and this reaction has done under vacuum then allowed to cool to room temperature, washed with methanol, and then washed with deionized water at 70-80°C. Finally, the product was dried under vacuum to absorb the acid from the substance. Increasing time of reaction does not have effect on reaction efficiency only purity of product get decreased. 2 g $\text{H}_2(\text{TCPP})$ poured into balloon was stirred under N_2 . After adding 25 mL DMSO and 4.3 g Ga(III), acetate $4\text{H}_2\text{O}$, the mixture was refluxed at 130°C for 6 h, it was allowed to cool for 6 h after stopping N_2 to achieve separation, the solvent substance was vacuumed. The solid produce was washed in deionized water to separate and extract, and for additional purification the substance was next moved through a silica column with DMF and dried under vacuum at 120°C [51]. 2 g Fe_3O_4 / APTES nanoparticles with 0.247 g (DCHC) and 20 mL DMF were mixed under N_2 atmosphere; 0.458 g Ga(TCPP) mixed in 5 mL DMF were added and were refluxed under N_2 atmosphere at 140°C for 8 h. A high flow of N_2 was used to prevent decomposition of DMF during the reaction. The product was washed with DMF, DCM and methanol after reaction, until porphyrin with an electrostatic bond mixed with Fe_3O_4 / APTES exited quickly and the product remained at 80°C for 24 h, to be dried and extra solvent could be exhausted. DCHC was used as a catalyst for making the bond between COOH groups and NH_2 groups on the magnetite surface (Figure 2).

Sorption procedure: Batch-sorption were performed by mixing 10 mL Fluoride aqueous solution (10.0 mg L^{-1}) and various amounts (10-100 mg) dry powder of the Fe_3O_4 / APTES / Ga (TCPP) nanoparticles in using a sonication method for 30 min. then, the

sorbent was separated before measurement, the Fluoride removal efficiency was calculated according to Equation (1) in the text:

Fluoride removal efficiency (%)= $(C_0-C_f)/C_0 \times 100$ (1)

where C_0 and C_f are the initial and final concentration of the fluoride ion, before and after the sorption, respectively.



Results and Discussion

Characterization of nanomagnets

XRD spectrum: Structure and particle size were determined by XRD, the crystalline structure of synthesized nanoparticles is shown in the XRD analysis. The Joint Committee on Powder Diffraction Standards (JCPDS) reference pattern of magnetite (No.19-629) is included for comparison. The XRD spectrum shows diffractions and the fact that Fe_3O_4 is available in the spinel phase. Existing peaks in (220), (311), (400), (511) and (411) indicate cubic spinel phases of Fe_3O_4 . Fe_3O_4 is completely pure due to lack of any peaks in 20-30 For $\gamma-Fe_2O_3$ [52], observation of peaks in the first section indicates synthesized magnetite; the second section shows modification of the magnetite surface and the third section is related to functionalization of the magnetite. All of the peaks in three parts are similar, indicating the stability of the crystalline phase of the nanoparticles. The results of the XRD analysis from the synthesized substance indicate the diameter of particles to be equal to 45 nm in histogram.

FT-IR spectrum: This method was employed for determining functional groups this is represented in Table 1. This Table displays all of the parts for three curve (a), (b), (c). Curve (a) relates to Fe_3O_4 and curve (b) associates to Fe_3O_4 /APTES and curve (c) links to Fe_3O_4 /APTES/Ga (TCPP) (Figure 3).

Functional Groups	Wavenumber (cm ⁻¹)
Fe-O	580 cm ⁻¹
Si-O-Si	1115 cm ⁻¹
Si-OH	1030 cm ⁻¹
N-H stretching vibration	1640 cm ⁻¹ and 3445 cm ⁻¹
C=H stretching vibration	2862 cm ⁻¹ and 2930 cm ⁻¹
O=C-N amide group	1687 cm ⁻¹
C=C stretching vibration	1500 cm ⁻¹

Table 1: Indicating Functional Groups with Wavenumber (cm⁻¹).

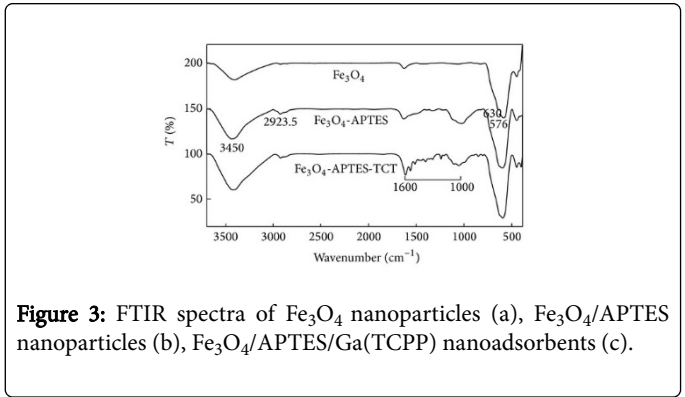


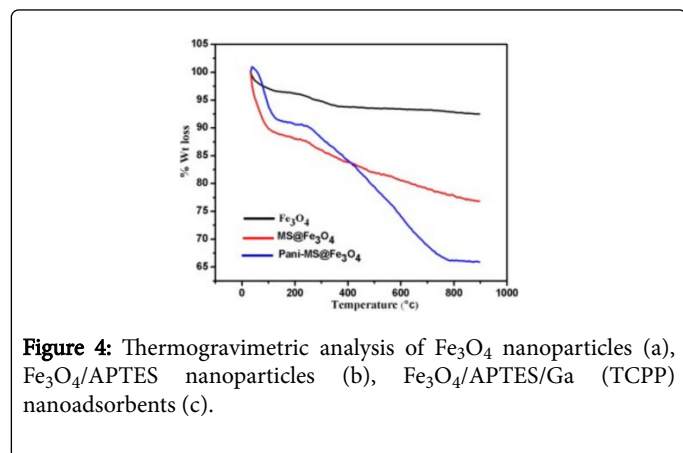
Figure 3: FTIR spectra of Fe_3O_4 nanoparticles (a), Fe_3O_4 /APTES nanoparticles (b), Fe_3O_4 /APTES/Ga(TCPP) nanoadsorbents (c).

Thermogravimetric Analysis (TGA): Consider (Table 2), in this study to identify the number of modifiers quantitatively, a TGA thermogram was used. This thermogram was illustrated in weight loss Curve (a) relates to Fe_3O_4 , curve (b) associates to Fe_3O_4 /APTES and curve (c) links to Fe_3O_4 /APTES/Ga (TCPP) [46]. These data show that Fe_3O_4 /APTES nanoparticles have been modified with metalloporphyrin successfully using a chemical bond. Two important advantages for this experiment were observed. The first advantage was the formation of a chemical bond of metalloporphyrin ligand on the Fe_3O_4 /APTES nanoparticles, which was physical sorption; the second advantage was the use of an external magnetic field for regenerating the nanosorbant [53-55] this is illustrated in Figure 4.

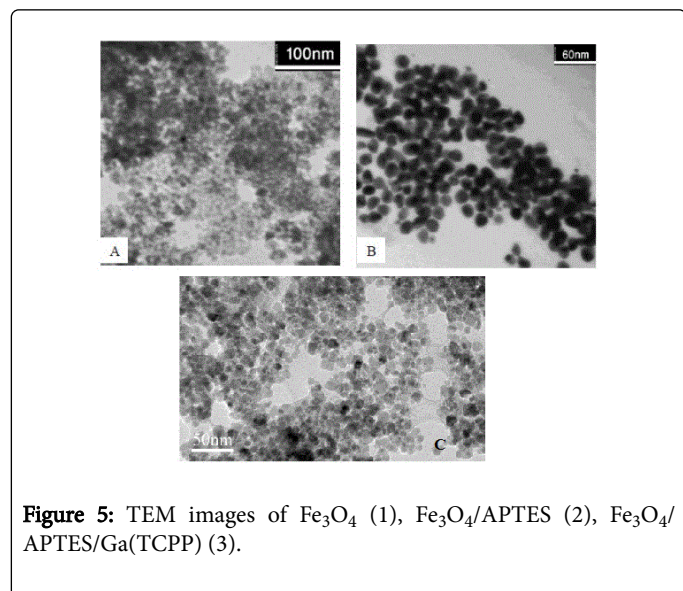
Fe_3O_4	Stages	
a) Fe_3O_4	Weight loss of 1%	Oxidation Fe_3O_4 to Fe_2O_3
b) Fe_3O_4 /APTES	The first weight loss stage was 2%	Decomposition of APTES
	The second weight loss stage was 7%	Gathering silanol group (OH)
C) Fe_3O_4 /APTES/Ga (TCPP)	The first weight loss stage was 4%	Missing magnetite humidity
	The second weight loss stage was 6%	Decreasing the APTES layer

	The third weight loss stage was 7%	Decomposition of metalloporphyrin
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Table 2: a) Fe_3O_4 with one stage of weight loss, b) Fe_3O_4 /APTES with two stages of weight loss, c) Fe_3O_4 /APTES/Ga (TCPP) with three stages of weight loss.



TEM analysis: As shown in Figure 5(a) The black powder is Fe_3O_4 , that is completely clear and pure due to lack of any peaks in 20-30 For $\gamma\text{-Fe}_2\text{O}_3$ [8] in XRD pattern and Figure 5(b) shows the APTES causes the some of the black powders Fe_3O_4 gather together and Figure 5(c) depicts the core of dark magnetic nanoparticles is coated with a gray shell. It is shown that magnetic nanoparticles are covered by Ga(TCPP); the magnetic nanoparticles have a spherical shape and the majority of particles are in the size range 40-50 nm.

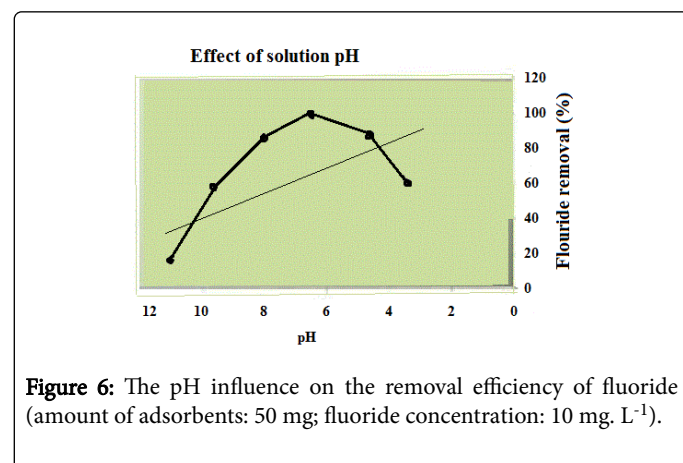


Optimization of the parameters

Effect of solution pH: Power of Hydrogen (pH) is one of the most important factors in substance adsorption in that suitable pH can raise yield sorption. In this study, pH values in the range between 3-11 were considered to measure the pH. Based on results, the highest fluoride removal percentage with Fe_3O_4 /APTES/Ga (TCPP) nanosorbent in 95%-96% was seen at pH values in the range between 5-7. At lower pH

values the fluoride removal decreased, since the metalloporphyrin complex experienced only minor demetalization during contact with the low pH solution and at higher pH values the fluoride removal also decreased due to the existence of competition between the hydroxyl groups and fluoride ions on active sites of the sorbent [37,42,56,57] are presented in Figure 6.

Effect of the adsorbent dosage: It was seen that by raising the amount of adsorbent, fluoride removal was increased, reaching a steady value for 50 mg adsorbent dosage. Using various amounts (10-100 mg) of Fe_3O_4 /APTES/Ga (TCPP) nanosorbents in fluoride removal achieved 96%. Thus, for all the experiments in this study, the adsorbent dosage was 50 mg [58,59] (Figure 7).



Effect of the contact time: The time of fluoride removal was considered for the Fe_3O_4 /APTES/Ga (TCPP) nanosorbent in 0-60 minutes, which determined the effect of the contact time on the fluoride removal with nanosorbent. It was seen that fluoride removal from the Fe_3O_4 /APTES/Ga (TCPP) nanosorbents was rapid at 30 min, so there was not a higher percent in the removal of fluoride are presented in Figure 8, in contact time 30 min fluoride removal achieved 93%.

Removal of fluoride from spiked sample: The solution of fluoride with 10 mg/L concentration was spiked with deionized water. The adsorption experiment was done on 10 mL of this spiked sample under optimal condition (contact time, 30 min, pH 5.5, and Fe_3O_4 /APTES/Ga (TCPP) nanosorbent 50 mg). It was seen that fluoride content decreased from 10.0 mg/L to 0.3 mg/L (97.0 ± 0.7%).

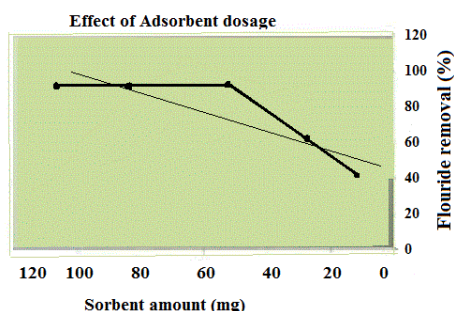


Figure 7: Effect of nanoadsorbent dosage on the on-fluoride removal (fluoride concentration: 10 mg L^{-1} , pH of the solution: 5.5, contact time: 30 min).

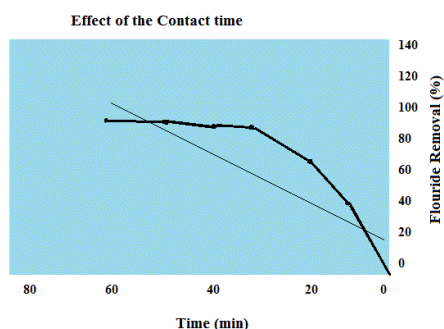


Figure 8: Effect of contact time on the adsorption of fluoride by modified nanoparticles (pH of the solution=5.5; amount of adsorbent=50 mg; fluoride concentration= 10 mg. L^{-1}).

Effect of co-existing anions: Wastewater has many ions with different concentration which depends on their source. The presence of these ions may have an effect on the extraction efficiency of $\text{Fe}_3\text{O}_4/\text{APTES}/\text{Ga}$ (TCPP) nanosorbents because of competition for binding with active sites between fluoride ions and foreign ions. The presence of other ions may interfere in the indicated fluoride adsorption; for this reason, the effect of foreign ions in fluoride removal has been studied. Some of the co-existing anion SO_4^{2-} , NO_3^- , NO_2^- , Cl^- were investigated (the concentration of every anion was 10 mg/L) Because the tolerance range was defined at $\pm 5\%$, and the presence of anions had no notable effect on the fluoride removal efficiency in solution, it is possible for Ga (III) porphyrin complexes to form selective bonds with fluoride [60].

Regeneration of the used sorbent: Regeneration is very important in the development of an economic process. In this study, NaOH solution was selected for regeneration $\text{Fe}_3\text{O}_4/\text{APTES}/\text{Ga}$ (TCPP) nanosorbent from fluoride ions. The concentration obtained by using NaOH 10.0 mM solution after optimizing was 98%, indicated the highest recovery of fluoride ions from the adsorbents (Figure 9).

Removal of fluoride from real sample: A sample was selected from wastewater released by the glass industry (Kaveh glass industry group, Saveh, Iran) as can be seen from Table 3. A comparison was performed between fluoride removal efficiency of the proposed nanosorbent with the commercially available activated alumina (KHD-12 from Sumitomo Chemical Co Ltd, Japan). The results show the untreated

wastewater had fluoride concentration of $35.5 \pm 1.2 \text{ mg/L}$. It has been found that the proposed nanosorbent efficiently removes fluoride from wastewater and decreases the fluoride concentration to $2.8 \pm 0.03 \text{ mg/L}$ which was fairly the same as the commercial one ($2.5 \pm 0.08 \text{ mg/L}$), showing that the proposed nanosorbent can be exploited for defluoridation in real field condition.

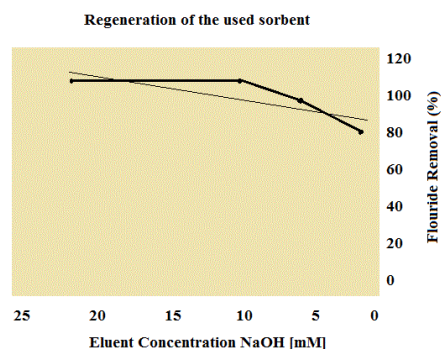


Figure 9: The effect of NaOH concentration as desorbing eluent on the recovery percentage of fluoride ion.

Parameter	Value (mgL^{-1})
Chloride	228 ± 19
Fluoride	35.5 ± 1.2
Nitrate	8.52 ± 0.6
Sulphate	108 ± 15
Bromide	0.06 ± 0.004
Nitrite	0.12 ± 0.006

Table 3: The characteristics of glass industrial wastewater used in this study.

Reusability of the recycled sorbent: Regeneration and recovery can be considered important factors for evaluating the effect of efficiency. After the first regeneration, this process was continued for five cycles. In this study, the fluoride removal efficiency of the $\text{Fe}_3\text{O}_4/\text{APTES}/\text{Ga}$ (TCPP) nanoadsorbents from five cycles were 97.2%, 96.8%, 95.5%, 94.3%, and 87.0% respectively. It was observed that modified nanoparticles can regenerate and be used more than once, providing an economic advantage for this experiment.

Conclusion

This technique has been done in controlling the concentration of fluoride in water problems at mining operation that it effects on the environment that it is one of the remarkable targets mine water and the environment. Fluoride pollution of the industrial wastewaters is one of the significant problems worldwide imposing a serious threat to human health. Magnetite nanoparticles were modified by grafting Ga(TCPP) groups at the surface of $\text{Fe}_3\text{O}_4/\text{APTES}$, and their ability to extract fluoride ions from aqueous solutions was evaluated. This study indicates that magnetite nanoparticles show good efficiency in removing fluoride ions and they can be selective for target anions. The results also indicate that magnetite nanoparticles can be used

interchangeably as magnetic Ionics, with the ability to separate fluoride ions from other liquid solution. The pH efficiency and presence of other ions were also investigated for removal efficiency. Two main advantages of the nanomaterial produced are ease of removal from the magnetic field and the possibility for easy regeneration after washing the fluoride ion with alkaline solution. Regeneration of the fluoride Ion and the feasibility for reuse are other advantages of this method.

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